Please check the examination details below before entering your candidate information

Candidate surname	Other names

Centre Number | Candidate Number

Pearson Edexcel International Advanced Level

Time 1 hour 30 minutes

Paper reference **WCH12/01**

Chemistry

International Advanced Subsidiary/Advanced Level
UNIT 2: Energetics, Group Chemistry, Halogenoalkanes and Alcohols

You must have:
Scientific calculator, Data Booklet, ruler

Total Marks

Instructions

- Use **black** ink or ball-point pen.
- If pencil is used for diagrams/sketches/graphs it must be dark (HB or B).
- **Fill in the boxes** at the top of this page with your name, centre number and candidate number.
- Answer **all** questions.
- Answer the questions in the spaces provided – *there may be more space than you need*.

Information

- The total mark for this paper is 80.
- The marks for **each** question are shown in brackets – *use this as a guide as to how much time to spend on each question*.
- In the question marked with an **asterisk** (*), marks will be awarded for your ability to structure your answer logically, showing how the points that you make are related or follow on from each other where appropriate.
- A Periodic Table is printed on the back cover of this paper.

Advice

- Read each question carefully before you start to answer it.
- Show all your working in calculations and include units where appropriate.
- Try to answer every question.
- Check your answers if you have time at the end.

Turn over ▶

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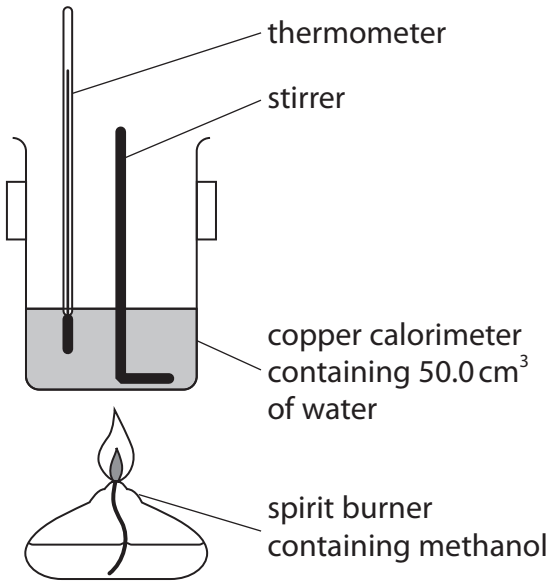
SECTION A

Answer ALL the questions in this section.

You should aim to spend no more than 20 minutes on this section.

For each question, select one answer from A to D and put a cross in the box ☒. If you change your mind, put a line through the box ☒ and then mark your new answer with a cross ☒.

1 A student measures the enthalpy change of combustion, $\Delta_c H$, of methanol, CH_3OH, using the apparatus shown.



After burning 0.20 g of methanol, the temperature of the water increases by 16.0 °C.

(a) The measurement uncertainty in the thermometer used in the experiment is 0.5 °C for each reading.

What is the percentage uncertainty in the temperature change of 16.0 °C?

(1)

☒ **A** 1.6%

☒ **B** 3.1%

☒ **C** 6.3%

☒ **D** 12.5%

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(b) The student repeats the experiment but burns 0.30 g of methanol and uses 75.0 cm^3 of water in the copper calorimeter.

What is the expected temperature change in this repeat experiment?

(1)

☒ **A** 7.1 °C

☒ **B** 10.7 °C

☒ **C** 16.0 °C

☒ **D** 36.0 °C

(c) The student's calculated enthalpy change of combustion of methanol is **more** exothermic than a data book value.

What is a possible reason for this?

(1)

☒ **A** heat loss to the surroundings

☒ **B** incomplete combustion of methanol

☒ **C** evaporation of methanol from the wick of the spirit burner

☒ **D** use of the molar mass of ethanol, C_2H_5OH, in the calculation

(Total for Question 1 = 3 marks)

2 Which equation represents the standard enthalpy change of atomisation, $\Delta_{at}H^{\ominus}$, of bromine?

☒ **A** $\frac{1}{2}Br_2(l) \rightarrow Br(g)$

☒ **B** $\frac{1}{2}Br_2(g) \rightarrow Br(g)$

☒ **C** $Br_2(l) \rightarrow 2Br(g)$

☒ **D** $Br_2(g) \rightarrow 2Br(g)$

(Total for Question 2 = 1 mark)

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3 The enthalpy change of reaction, $\Delta_r H$, for the equation shown can be calculated using bond enthalpy data.

$$\tfrac{1}{2}H_2(g) + \tfrac{1}{2}Cl_2(g) \rightarrow HCl(g)$$

Bond	Bond enthalpy / kJ mol^{-1}
H—H	436
Cl—Cl	242
H—Cl	431

The expression that should be used in the calculation is

☒ **A** $(0.5 \times 436 + 0.5 \times 242) - 431$

☒ **B** $(2 \times 431) - (436 + 242)$

☒ **C** $431 - (0.5 \times 436 + 0.5 \times 242)$

☒ **D** $(436 + 242) - (2 \times 431)$

(Total for Question 3 = 1 mark)

4 Which compound has London forces as the **only** intermolecular force?

☒ **A** HF

☒ **B** OF_2

☒ **C** PF_3

☒ **D** CF_4

(Total for Question 4 = 1 mark)

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5 Which compound has intermolecular hydrogen bonding?

- ☒ **A** $(CH_3)_3N$
- ☒ **B** $(CH_3)_3CF$
- ☒ **C** $(CH_3)_3COH$
- ☒ **D** $(CH_3)_3CCHO$

(Total for Question 5 = 1 mark)

6 Which sequence shows the hydrogen halides in order of **decreasing** boiling temperature?

- ☒ **A** HF > HCl > HBr > HI
- ☒ **B** HF > HI > HBr > HCl
- ☒ **C** HI > HF > HCl > HBr
- ☒ **D** HI > HBr > HCl > HF

(Total for Question 6 = 1 mark)

7 Which ion contains vanadium with an oxidation number of **+4**?

- ☒ **A** VO^{2+}
- ☒ **B** VO_2^+
- ☒ **C** VO_3^-
- ☒ **D** VO_4^{3-}

(Total for Question 7 = 1 mark)

8 What is the formula of potassium manganate(VI)?

- ☒ **A** $KMnO_4$
- ☒ **B** K_2MnO_4
- ☒ **C** K_3MnO_4
- ☒ **D** $K_6Mn_2O_6$

(Total for Question 8 = 1 mark)

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9 Compound **Q** produces

- a red colour in a flame test
- a white precipitate when aqueous potassium sulfate is added to a solution of the compound.

What is compound **Q**?

☒ **A** LiCl

☒ **B** $NaNO_3$

☒ **C** $Sr(NO_3)_2$

☒ **D** $BaCl_2$

(Total for Question 9 = 1 mark)

10 Which reaction produces more than one product?

☒ **A** $Mg + O_2 \rightarrow$

☒ **B** $Ca + Cl_2 \rightarrow$

☒ **C** $Sr + H_2O \rightarrow$

☒ **D** $BaO + H_2O \rightarrow$

(Total for Question 10 = 1 mark)

11 Which equation shows a redox reaction that would **not** be expected to occur, based on the trend in reactivity of the halogens?

☒ **A** $I_2(aq) + 2At^-(aq) \rightarrow At_2(aq) + 2I^-(aq)$

☒ **B** $2Br^-(aq) + Cl_2(aq) \rightarrow 2Cl^-(aq) + Br_2(aq)$

☒ **C** $Cl_2(aq) + 2I^-(aq) \rightarrow I_2(aq) + 2Cl^-(aq)$

☒ **D** $2F^-(aq) + At_2(aq) \rightarrow 2At^-(aq) + F_2(aq)$

(Total for Question 11 = 1 mark)

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12 A fixed amount of concentrated H_2SO_4 is reacted separately with an excess of four solid potassium halides.

In which reaction would the greatest number of moles of halide be oxidised?

- ☒ **A** $2KF(s) + H_2SO_4(aq) \rightarrow K_2SO_4(aq) + 2HF(g)$
- ☒ **B** $KCl(s) + H_2SO_4(aq) \rightarrow KHSO_4(aq) + HCl(g)$
- ☒ **C** $2KBr(s) + 3H_2SO_4(aq) \rightarrow 2KHSO_4(aq) + Br_2(l) + SO_2(g) + 2H_2O(l)$
- ☒ **D** $8KI(s) + 9H_2SO_4(aq) \rightarrow 8KHSO_4(aq) + 4I_2(s) + H_2S(g) + 4H_2O(l)$

(Total for Question 12 = 1 mark)

13 Silver nitrate in aqueous ethanol is added separately to four halogenoalkanes.

Which would form a silver halide precipitate in the **shortest** time?

- ☒ **A** $(CH_3)_3CI$
- ☒ **B** $(CH_3)_3CCl$
- ☒ **C** $CH_3CH_2CH_2CH_2I$
- ☒ **D** $CH_3CH_2CH_2CH_2Cl$

(Total for Question 13 = 1 mark)

14 $CH_3CH_2CHBrCH_2CH_2CH_3$ is heated with ethanolic potassium hydroxide.

How many **alkene** products are possible?

- ☒ **A** one
- ☒ **B** two
- ☒ **C** three
- ☒ **D** four

(Total for Question 14 = 1 mark)

15 Which compound is **least** likely to have a prominent peak at $m/z = 43$ in its mass spectrum?

- ☒ **A** $CH_3COCH_2CH_3$
- ☒ **B** $CH_3CH_2NHCH_3$
- ☒ **C** $CH_3CH(CH_3)_2$
- ☒ **D** $CH_3CH_2CH_2CH_3$

(Total for Question 15 = 1 mark)

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16 Which compound has peaks at 3415 cm^{-1} **and** 2250 cm^{-1} in its infrared spectrum? Refer to the Data Booklet.

- ☒ **A** $H_2NCH_2CH_2C{\equiv}N$
- ☒ **B** $CH_3CH_2CH_2COOH$
- ☒ **C** $ClCH_2CH_2C{\equiv}CH$
- ☒ **D** $HOCH_2CH_2CH{=}CH_2$

(Total for Question 16 = 1 mark)

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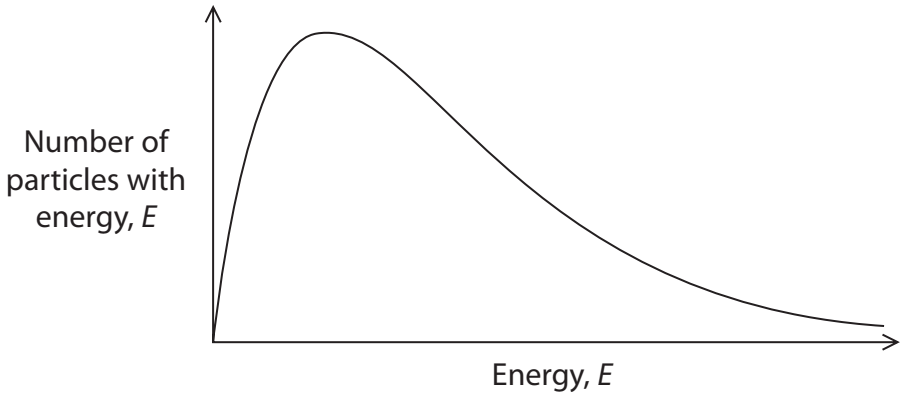
17 The distribution of molecular energies for a sample of gas in a sealed container is shown.



(a) Why does the distribution of energies start at the origin, (0,0)?

(1)

- ☒ **A** some molecules have no energy
- ☒ **B** all molecules possess some energy
- ☒ **C** the temperature is 0 K
- ☒ **D** some molecules do not have enough energy to react

(b) Some of the gas is removed and then the container is resealed and the gas is cooled.

How does the new distribution of molecular energies compare to the original sample?

(1)

	Area under the curve	Position of peak
☒ **A**	does not change	shifts to the left
☒ **B**	decreases	shifts to the left
☒ **C**	does not change	shifts to the right
☒ **D**	decreases	shifts to the right

(Total for Question 17 = 2 marks)

TOTAL FOR SECTION A = 20 MARKS

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SECTION B

Answer ALL the questions. Write your answers in the spaces provided.

18 Calcium hypochlorite, $Ca(ClO)_2$, is used for water treatment in swimming pools. It is produced in the reaction between $Ca(OH)_2$ and Cl_2.

$$2Ca(OH)_2(aq) + 2Cl_2(g) \rightarrow Ca(ClO)_2(aq) + CaCl_2(aq) + 2H_2O(l)$$

(a) State the type of reaction occurring in the production of $Ca(ClO)_2$. Justify your answer using oxidation numbers.

(3)

(b) Calculate the percentage atom economy by mass for the production of $Ca(ClO)_2$ in this reaction.

(2)

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(c) A swimming pool has the dimensions 50 m × 25 m × 2.0 m.
The water in this swimming pool has a $Ca(ClO)_2$ concentration of 4.2 mg dm^{-3}.

(i) Calculate the mass, **in kg**, of $Ca(ClO)_2$ required to treat the water needed to completely fill this swimming pool.

(3)

(ii) Calculate the volume of Cl_2, at room temperature and pressure, needed to make the mass of $Ca(ClO)_2$ calculated in (c)(i).

(3)

(Total for Question 18 = 11 marks)

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19 This question is about alcohols with the molecular formula $C_6H_{14}O$.

(a) Draw the **skeletal** formula of each of the three **tertiary** alcohols with the formula $C_6H_{14}O$.

(3)

tertiary alcohol 1

tertiary alcohol 2

tertiary alcohol 3

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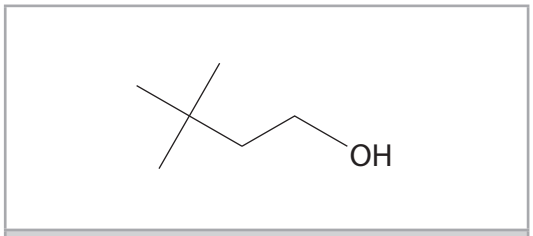
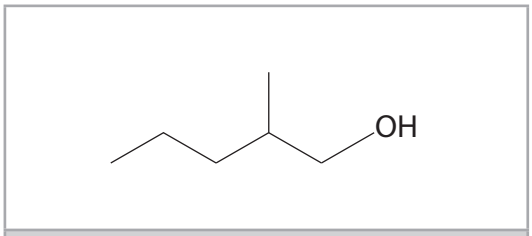
(b) Two primary $C_6H_{14}O$ alcohols, **A** and **B**, are shown.



alcohol **A**



alcohol **B**

(i) Give the IUPAC name of alcohol **A**.

(1)

(ii) Explain why alcohol **B** has a higher boiling temperature than alcohol **A**.

(2)

(iii) Explain why alcohol **B** is completely soluble in ethanol but only slightly soluble in water.
A detailed description of the forces involved is **not** required.

(3)

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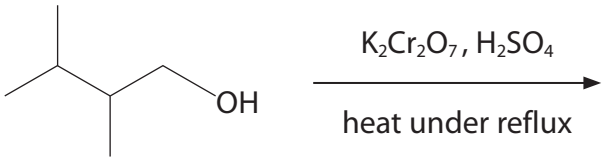
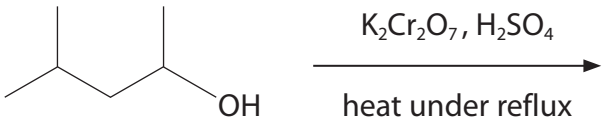
(c) Give the structure of the **organic** product of each reaction shown.

(3)

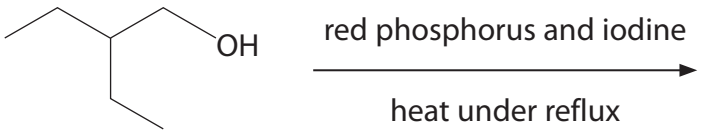
Reaction **1**



Reaction **2**



Reaction **3**



(Total for Question 19 = 12 marks)

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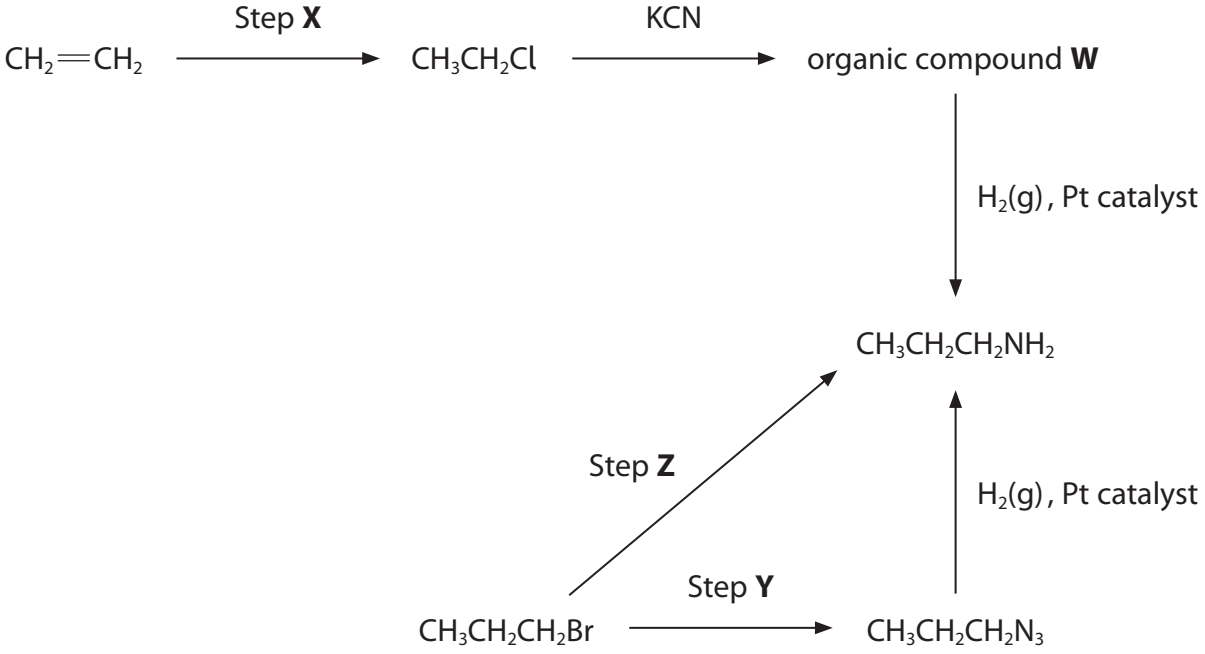
20 This question is about the synthesis of propylamine, $CH_3CH_2CH_2NH_2$.

Three routes for the synthesis of $CH_3CH_2CH_2NH_2$ are shown.



(a) Identify, by name or formula, the reagent used in Step **X**.

(1)

.....

(b) Give the structure of organic compound **W**.

(1)

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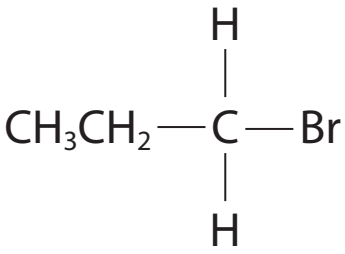
(c) The reagent used in Step **Y** is potassium azide, KN_3.
This is a source of the azide ion, N_3^-, which acts as a nucleophile.

(i) Complete a possible dot-and-cross diagram for N_3^-.
Show outer electrons only.

(2)

(ii) Complete the mechanism for Step **Y**.
Include curly arrows, and any relevant lone pairs and dipoles.

(3)



N_3^-

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(d) Step **Z** is carried out by reacting ammonia with 1-bromopropane.

(i) Give the conditions for this reaction.

(2)

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(ii) Suggest why the yield of $CH_3CH_2CH_2NH_2$ obtained using Step **Z** is low.

(1)

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(Total for Question 20 = 10 marks)

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***21** Discuss some aspects of the thermal stability of the anhydrous nitrates of the elements in Groups 1 and 2 of the Periodic Table.

In your answer you should

- explain the trend in thermal stability of the **Group 2** nitrates
- describe any differences in the products of thermal decomposition of the **Group 1** nitrates
- give equations for the thermal decomposition of sodium nitrate and of magnesium nitrate.

(6)

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(Total for Question 21 = 6 marks)

TOTAL FOR SECTION B = 39 MARKS

SECTION C

Answer ALL the questions. Write your answers in the spaces provided.

22 Ammonium nitrate, NH_4NO_3, is used in the manufacture of fertilisers and explosives.
It is produced on a large scale using only methane, water and air.
The process has four stages.

(a) The first two reactions in Stage **1** involve the production of hydrogen.

At temperature T_1, methane reacts with excess steam to give hydrogen.

$$CH_4(g) + H_2O(g) \rightleftharpoons 3H_2(g) + CO(g) \qquad \Delta H = +206\,kJ\,mol^{-1}$$

At a different temperature, T_2, the carbon monoxide reacts with more steam.

$$CO(g) + H_2O(g) \rightleftharpoons H_2(g) + CO_2(g) \qquad \Delta H = -42\,kJ\,mol^{-1}$$

(i) Give the reason why excess steam is used in the first reaction.

(1)

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(ii) Predict which of T_1 and T_2 is the **higher** temperature.
Justify your answer.

(1)

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(iii) Derive the **overall** equation for the production of H_2 in Stage **1**.
State symbols are not required.

(1)

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(b) The third reaction in Stage **1** involves the removal of carbon dioxide, using an aqueous solution of N-methyldiethanolamine, $CH_3N(CH_2CH_2OH)_2$.

$$CO_2(g) + H_2O(l) + 2CH_3N(CH_2CH_2OH)_2(aq) \rightleftharpoons (CH_3NH(CH_2CH_2OH)_2)_2CO_3(aq)$$

(i) Suggest **one** reason why CO_2 is removed.

(1)

(ii) Name the type of reaction occurring.

(1)

(iii) Draw the **displayed** formula of N-methyldiethanolamine, $CH_3N(CH_2CH_2OH)_2$.

(1)

(c) In Stage **2**, the hydrogen from Stage **1** reacts with nitrogen (from the air) to produce ammonia. The conditions for this reaction are:

- a temperature of 700 K
- a pressure in the range 100–200 atm
- an iron catalyst

$$N_2(g) + 3H_2(g) \rightleftharpoons 2NH_3(g) \qquad \Delta H = -92\,kJ\,mol^{-1}$$

Give **one** advantage and **one** disadvantage of using a pressure of 200 atm, compared to a pressure of 100 atm, in Stage **2**.

(2)

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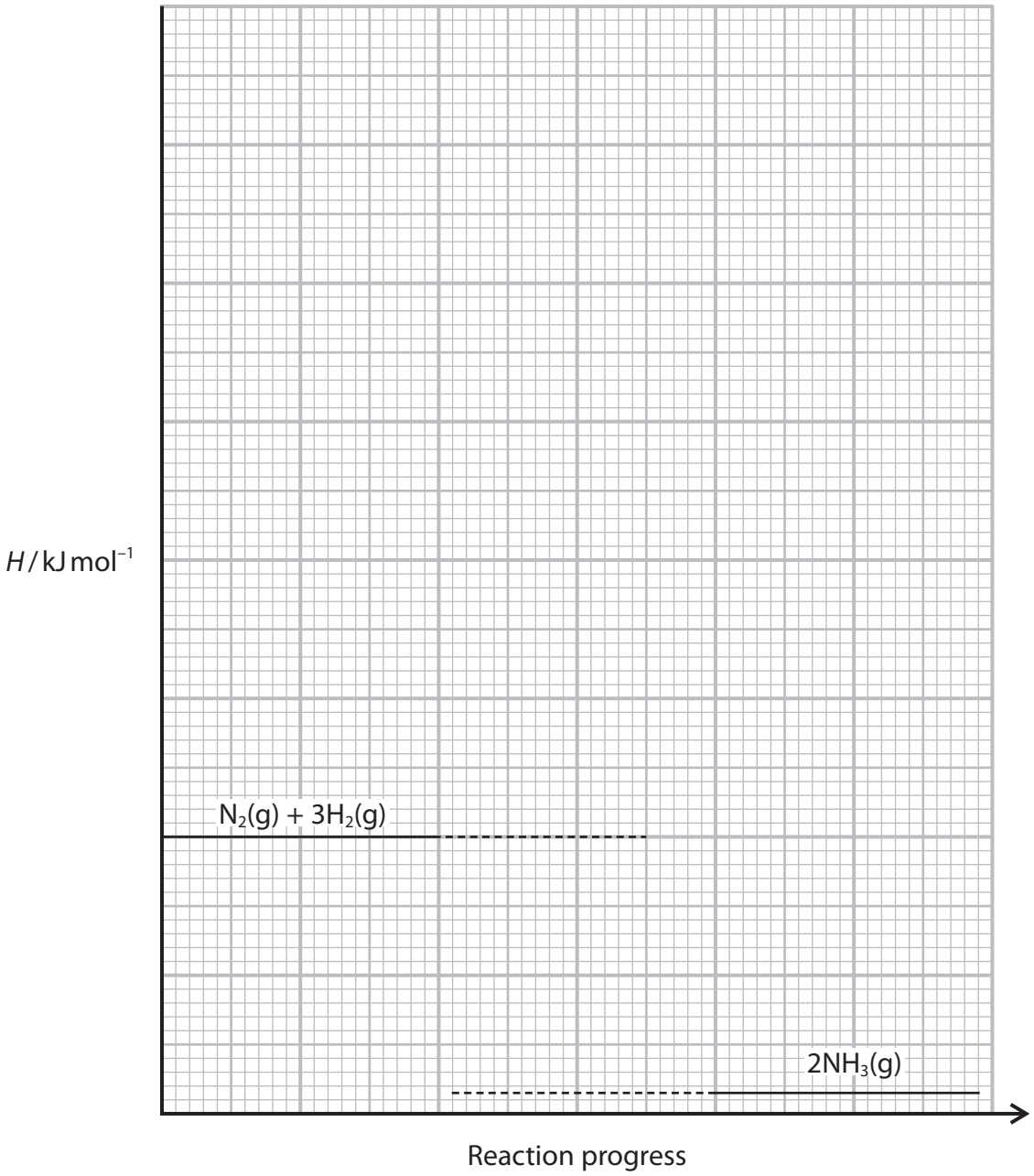
(d) The reaction in Stage **2** has an activation energy, $E_{cat} = +70\,kJ\,mol^{-1}$.

The **uncatalysed** reaction between N_2 and H_2 has an activation energy, $E_a = +290\,kJ\,mol^{-1}$.

(i) Complete the profile for the catalysed and uncatalysed reactions.
Label the activation energies and the enthalpy change of reaction, ΔH.

Your diagram **must** match the scale shown for the production of NH_3.

(3)



(ii) Suggest why the use of the catalyst makes Stage **2** more sustainable.

(1)

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(e) In Stage **3**, nitrogen monoxide, NO, is produced in the reaction between NH_3 (from Stage **2**) and O_2 (from the air). The conditions used are a temperature of 1100 K in the presence of a platinum-rhodium catalyst.

$$4NH_3(g) + 5O_2(g) \rightleftharpoons 4NO(g) + 6H_2O(g) \qquad \Delta H = -905\ kJ\ mol^{-1}$$

(i) Give **one** reason why a high temperature is needed in this reaction.

(1)

(ii) Suggest why only a small amount of energy is used to maintain the temperature at 1100 K.

(1)

(f) The NO from the first reaction in Stage **3** is cooled and then converted to nitrogen dioxide, NO_2, by reaction with more O_2.

$$2NO(g) + O_2(g) \rightleftharpoons 2NO_2(g)$$

Nitric acid, $HNO_3(aq)$, is produced by the addition of water.

$$3NO_2(g) + H_2O(l) \rightarrow 2HNO_3(aq) + NO(g)$$

Explain how adding water in the second reaction affects the yield of NO_2 in the first reaction.

(2)

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(g) In Stage **4**, a solution of NH_4NO_3 is produced by reacting NH_3 (from Stage **2**) with HNO_3 (from Stage **3**).

$$NH_3(g) + HNO_3(aq) \rightarrow NH_4NO_3(aq)$$

Data

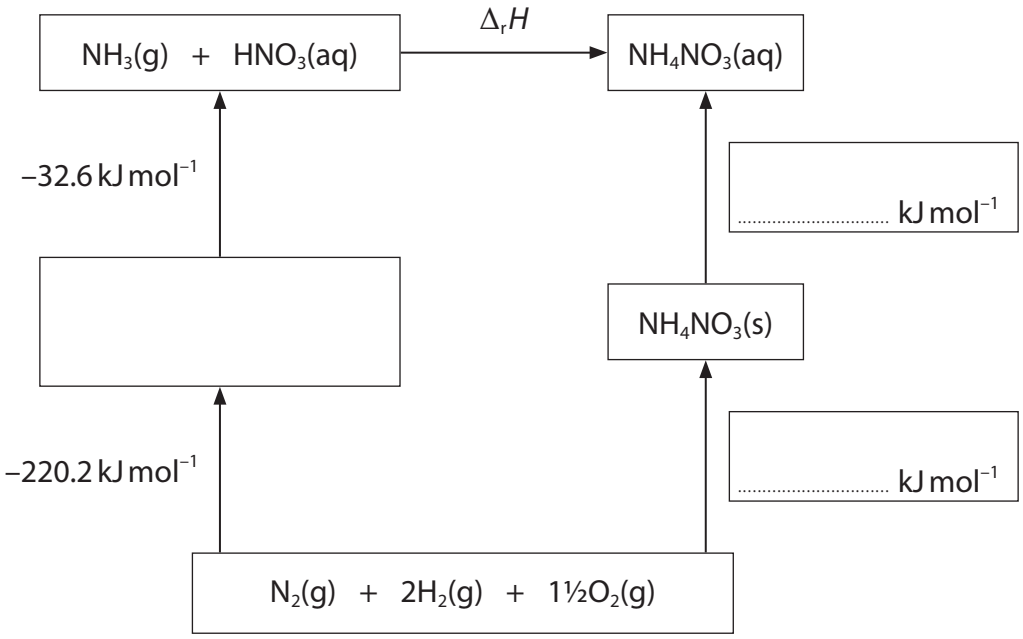
Species	$NH_3(g)$	$HNO_3(l)$	$NH_4NO_3(s)$
$\Delta_f H^{\ominus}$ / kJ mol^{-1}	–46.1	–174.1	–365.6

Equation	ΔH / kJ mol^{-1}
$HNO_3(l) + aq \rightarrow HNO_3(aq)$	–32.6
$NH_4NO_3(s) + aq \rightarrow NH_4NO_3(aq)$	+25.6

(i) Complete the enthalpy cycle.

(2)

Enthalpy cycle



(ii) Calculate the enthalpy change, $\Delta_r H$, in kJ mol^{-1}, for the reaction of $NH_3(g)$ with $HNO_3(aq)$.

(1)

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(h) Suggest **two** reasons why it is more profitable to carry out all four stages at the **same** site, instead of using different sites for each stage in the industrial production of ammonium nitrate.

(2)

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(Total for Question 22 = 21 marks)

TOTAL FOR SECTION C = 21 MARKS
TOTAL FOR PAPER = 80 MARKS

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The Periodic Table of Elements

1	2											3	4	5	6	7	0 (8)
																	(18)
							1.0 H hydrogen 1										4.0 He helium 2
(1)	(2)		Key									(13)	(14)	(15)	(16)	(17)	
6.9 Li lithium 3	9.0 Be beryllium 4		relative atomic mass atomic symbol name atomic (proton) number									10.8 B boron 5	12.0 C carbon 6	14.0 N nitrogen 7	16.0 O oxygen 8	19.0 F fluorine 9	20.2 Ne neon 10
23.0 Na sodium 11	24.3 Mg magnesium 12	(3)	(4)	(5)	(6)	(7)	(8)	(9)	(10)	(11)	(12)	27.0 Al aluminium 13	28.1 Si silicon 14	31.0 P phosphorus 15	32.1 S sulfur 16	35.5 Cl chlorine 17	39.9 Ar argon 18
39.1 K potassium 19	40.1 Ca calcium 20	45.0 Sc scandium 21	47.9 Ti titanium 22	50.9 V vanadium 23	52.0 Cr chromium 24	54.9 Mn manganese 25	55.8 Fe iron 26	58.9 Co cobalt 27	58.7 Ni nickel 28	63.5 Cu copper 29	65.4 Zn zinc 30	69.7 Ga gallium 31	72.6 Ge germanium 32	74.9 As arsenic 33	79.0 Se selenium 34	79.9 Br bromine 35	83.8 Kr krypton 36
85.5 Rb rubidium 37	87.6 Sr strontium 38	88.9 Y yttrium 39	91.2 Zr zirconium 40	92.9 Nb niobium 41	95.9 Mo molybdenum 42	[98] Tc technetium 43	101.1 Ru ruthenium 44	102.9 Rh rhodium 45	106.4 Pd palladium 46	107.9 Ag silver 47	112.4 Cd cadmium 48	114.8 In indium 49	118.7 Sn tin 50	121.8 Sb antimony 51	127.6 Te tellurium 52	126.9 I iodine 53	131.3 Xe xenon 54
132.9 Cs caesium 55	137.3 Ba barium 56	138.9 La* lanthanum 57	178.5 Hf hafnium 72	180.9 Ta tantalum 73	183.8 W tungsten 74	186.2 Re rhenium 75	190.2 Os osmium 76	192.2 Ir iridium 77	195.1 Pt platinum 78	197.0 Au gold 79	200.6 Hg mercury 80	204.4 Tl thallium 81	207.2 Pb lead 82	209.0 Bi bismuth 83	[209] Po polonium 84	[210] At astatine 85	[222] Rn radon 86
[223] Fr francium 87	[226] Ra radium 88	[227] Ac* actinium 89	[261] Rf rutherfordium 104	[262] Db dubnium 105	[266] Sg seaborgium 106	[264] Bh bohrium 107	[277] Hs hassium 108	[268] Mt meitnerium 109	[271] Ds darmstadtium 110	[272] Rg roentgenium 111							

Elements with atomic numbers 112-116 have been reported but not fully authenticated

* Lanthanide series

* Actinide series

140 Ce cerium 58	141 Pr praseodymium 59	144 Nd neodymium 60	[147] Pm promethium 61	150 Sm samarium 62	152 Eu europium 63	157 Gd gadolinium 64	159 Tb terbium 65	163 Dy dysprosium 66	165 Ho holmium 67	167 Er erbium 68	169 Tm thulium 69	173 Yb ytterbium 70	175 Lu lutetium 71
232 Th thorium 90	[231] Pa protactinium 91	238 U uranium 92	[237] Np neptunium 93	[242] Pu plutonium 94	[243] Am americium 95	[247] Cm curium 96	[245] Bk berkelium 97	[251] Cf californium 98	[254] Es einsteinium 99	[253] Fm fermium 100	[256] Md mendelevium 101	[254] No nobelium 102	[257] Lr lawrencium 103

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