



UNIVERSITY OF CAMBRIDGE INTERNATIONAL EXAMINATIONS
International General Certificate of Secondary Education

CHEMISTRY

0620/13

Paper 1 Multiple Choice

October/November 2012

45 Minutes

Additional Materials: Multiple Choice Answer Sheet
 Soft clean eraser
 Soft pencil (type B or HB is recommended)



READ THESE INSTRUCTIONS FIRST

Write in soft pencil.

Do not use staples, paper clips, highlighters, glue or correction fluid.

Write your name, Centre number and candidate number on the Answer Sheet in the spaces provided unless this has been done for you.

There are **forty** questions on this paper. Answer **all** questions. For each question there are four possible answers **A, B, C** and **D**.

Choose the **one** you consider correct and record your choice in **soft pencil** on the separate Answer Sheet.

Read the instructions on the Answer Sheet very carefully.

Each correct answer will score one mark. A mark will not be deducted for a wrong answer.

Any rough working should be done in this booklet.

A copy of the Periodic Table is printed on page 16.

You may use a calculator.

This document consists of **14** printed pages and **2** blank pages.



- 1 'Particles moving **very slowly** from an area of high concentration to an area of low concentration.'

Which process is being described above?

- A** a liquid being frozen
B a solid melting
C a substance diffusing through a liquid
D a substance diffusing through the air
- 2 Solid W melts at exactly 54 °C and boils at exactly 302 °C.

Solid X, when dissolved in water and examined using paper chromatography, shows a blue colour and a red colour.

Which row is correct?

	contains only one substance	contains more than one substance
A	W and X	–
B	W	X
C	X	W
D	–	W and X

- 3 Part of the instructions in an experiment reads as follows.

Quickly add 50 cm³ of acid.

What is the best piece of apparatus to use?

- A** a burette
B a conical flask
C a measuring cylinder
D a pipette

- 4 The nucleon number of an isotope of rubidium is 85.

How many protons, neutrons and electrons are present in an atom of this isotope?

	protons	neutrons	electrons
A	37	48	37
B	37	48	39
C	39	46	37
D	39	46	39

- 5 Which row gives the number of electrons in the outer electron shell of fluorine and of neon?

	${}^{19}_{9}\text{F}$	${}^{20}_{10}\text{Ne}$
A	7	8
B	7	10
C	9	8
D	9	10

- 6 In the molecules CH_4 , HCl and H_2O , which atoms use **all** of their outer shell electrons in bonding?

A C and Cl **B** C and H **C** Cl and H **D** H and O

- 7 The table shows the electronic structures of four atoms.

atom	electronic structure
W	2,1
X	2,7
Y	2,8,4
Z	2,8,8

Which two atoms combine to form an ionic compound?

A W and X **B** W and Y **C** X and Y **D** X and Z

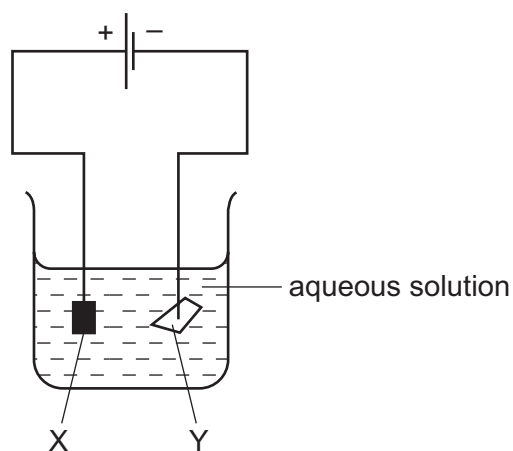
- 8 A compound has the formula $\text{CH}_3\text{CO}_2\text{H}$.

How should the relative molecular mass, M_r , of this compound be calculated?

- A $12 + 1 + 16$
 B $3(12 + 1) + 2(12 + 16) + 1$
 C $(4 \times 12) + (2 \times 1) + 16$
 D $(2 \times 12) + (4 \times 1) + (2 \times 16)$
- 9 In separate experiments, electricity was passed through concentrated aqueous sodium chloride and molten lead bromide.

What would happen in **both** experiments?

- A A halogen would be formed at the anode.
 B A metal would be formed at the cathode.
 C Hydrogen would be formed at the anode.
 D Hydrogen would be formed at the cathode.
- 10 The diagram shows an electrolysis experiment using metals X and Y as electrodes.

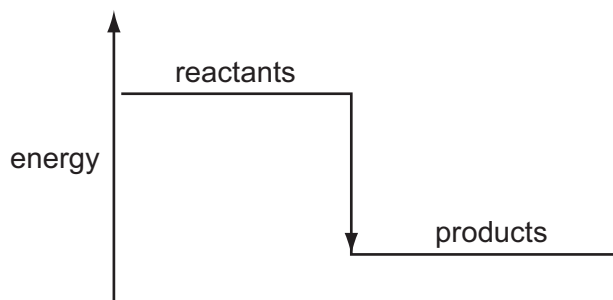


One of the metals becomes coated with copper.

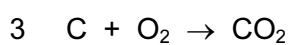
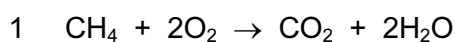
Which metal becomes coated and which aqueous solution is used?

	metal	aqueous solution
A	X	CrCl_3
B	X	CuCl_2
C	Y	CrCl_3
D	Y	CuCl_2

11 A diagram for the energy change during an exothermic reaction is shown.



For which reactions would this be an appropriate diagram?



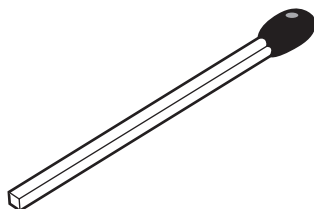
A none of them

B 1 and 2 only

C 2 and 3 only

D all of them

12 The diagram shows a match.



By striking the match, a chemical reaction takes place.

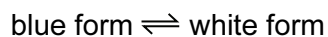
Which statements about the chemical reaction are correct?

	type of reaction	reason
A	endothermic	because energy is used to strike the match
B	endothermic	because energy is given out as the match burns
C	exothermic	because energy is used to strike the match
D	exothermic	because energy is given out as the match burns

13 Solid copper(II) sulfate can exist in two different forms, anhydrous and hydrated.

One of these forms is blue and the other is white.

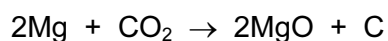
The change between these two forms is reversible.



Which is the blue form and how is the change from the blue form to the white form brought about?

	blue form	change to white form
A	anhydrous	add water
B	anhydrous	heat
C	hydrated	add water
D	hydrated	heat

14 Which statement describes what happens in the reaction shown?



- A** Carbon and magnesium are both oxidised.
- B** Carbon is oxidised and magnesium oxide is reduced.
- C** Magnesium is oxidised and carbon dioxide is reduced.
- D** Magnesium oxide and carbon dioxide are both reduced.

15 Which change does **not** increase the speed of reaction between zinc and hydrochloric acid?

- A** adding a catalyst
- B** decreasing the particle size of the zinc
- C** decreasing the temperature
- D** using more concentrated acid

16 Salts X and Y are separately dissolved in water.

Samples of the solutions obtained are separately tested with dilute hydrochloric acid and with aqueous sodium hydroxide.

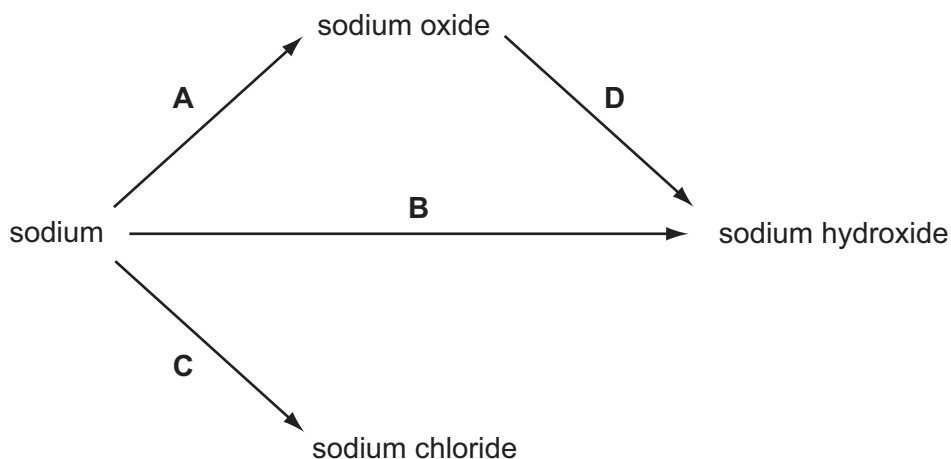
In two of the tests, a gaseous product is formed. No precipitate is formed in any of the tests.

What are salts X and Y?

	X	Y
A	AgNO ₃	BaSO ₄
B	BaSO ₄	Na ₂ CO ₃
C	Na ₂ CO ₃	NH ₄ Cl
D	NH ₄ Cl	AgNO ₃

17 Some reactions involving sodium are shown.

Which reaction does **not** involve the formation of a base?



18 Barium hydroxide is an alkali. It reacts with hydrochloric acid.

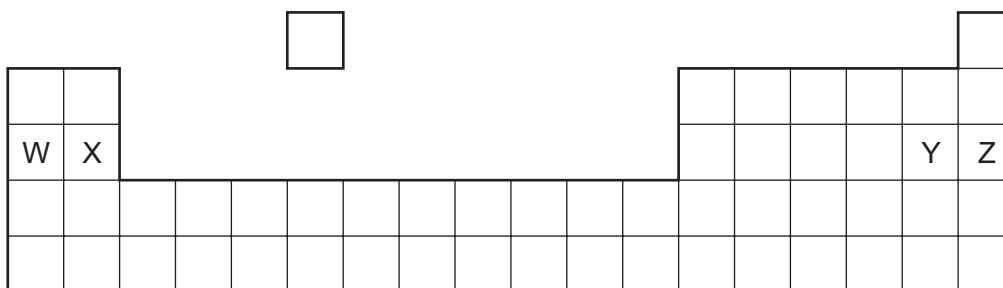
How does the pH of the hydrochloric acid change as an excess of aqueous barium hydroxide is added?

- A** The pH decreases from 14 and becomes constant at 7.
- B** The pH decreases from 14 to about 1.
- C** The pH increases from 1 and becomes constant at 7.
- D** The pH increases from 1 to about 14.

19 A compound is a salt if it

- A can neutralise an acid.
- B contains more than one element.
- C dissolves in water.
- D is formed when an acid reacts with a base.

20 The diagram shows an outline of the Periodic Table.

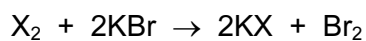


Which ionic compound could be formed?

- A W^+Y^-
- B W^+Z^-
- C X^+Y^-
- D X^+Z^-

21 Element X is in Group VII of the Periodic Table.

It reacts with aqueous potassium bromide as shown.



Which statements about X are correct?

	relative atomic mass	reactivity
A	greater than that of bromine	less reactive than bromine
B	greater than that of bromine	more reactive than bromine
C	less than that of bromine	less reactive than bromine
D	less than that of bromine	more reactive than bromine

22 The table gives information about four elements.

Which element is a transition metal?

	colour of element	electrical conductivity of element	colour of oxide
A	black	high	colourless
B	colourless	low	white
C	grey	high	red
D	yellow	low	colourless

23 Why are weather balloons filled with helium rather than hydrogen?

- A** Helium is found in air.
- B** Helium is less dense than hydrogen.
- C** Helium is more dense than hydrogen.
- D** Helium is unreactive.

24 The three statements below about aluminium are all correct.

- 1 Aluminium has an oxide ore.
- 2 Aluminium is extracted from its ore by electrolysis.
- 3 Aluminium is used to make cooking utensils.

In which of these statements can '**aluminium**' be replaced by '**iron**' and still be correct?

	1	2	3
A	✓	✓	✓
B	✓	✓	x
C	✓	x	✓
D	x	✓	✓

- 25 Statement 1: An aluminium alloy containing a small amount of copper is stronger than pure aluminium.

Statement 2: Alloys are made by mixing molten metals together and allowing them to cool.

Which of the following is true?

- A Both statements are correct and statement 2 explains statement 1.
- B Both statements are correct but statement 2 does not explain statement 1.
- C Statement 1 is correct but statement 2 is not correct.
- D Statement 2 is correct but statement 1 is not correct.
- 26 Iron is extracted from its ore in a Blast Furnace.
- Hematite, coke, limestone and hot air are added to the furnace.
- Which explanation is **not** correct?
- A Coke burns and produces a high temperature.
- B Hematite is the ore containing the iron as iron oxide.
- C Hot air provides the oxygen for the burning.
- D Limestone reduces the iron oxide to iron.
- 27 How does the reactivity of potassium compare with that of sodium and how does the reactivity of calcium compare with that of magnesium?

	reactivity of potassium and sodium	reactivity of calcium and magnesium
A	K greater than Na	Ca greater than Mg
B	K greater than Na	Mg greater than Ca
C	Na greater than K	Ca greater than Mg
D	Na greater than K	Mg greater than Ca

- 28 Untreated water can spread diseases such as cholera.

What can be added to drinking water to reduce the spread of such diseases?

- A carbon monoxide
- B chlorine
- C nitrogen monoxide
- D sulfur

29 Statement 1: The burning of fossil fuels containing sulfur is a cause of 'acid rain'.

Statement 2: Acid rain contains sulfur dioxide which is formed when sulfur compounds burn in the air.

Which of the following is true?

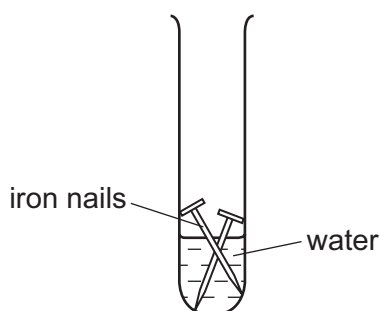
- A Both statements are correct and statement 2 explains statement 1.
- B Both statements are correct, but statement 2 does not explain statement 1.
- C Statement 1 is correct but statement 2 is incorrect.
- D Statement 2 is correct but statement 1 is incorrect.

30 Argon is a noble gas used to fill light bulbs.

What is the approximate percentage of argon in air?

- A 1%
- B 20%
- C 79%
- D 99%

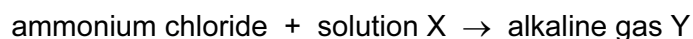
31 The experiment shown was set up to investigate the rusting of iron.



Rusting occurs in this experiment because of the presence of

- A hydrogen and oxygen.
- B nitrogen and oxygen.
- C nitrogen and water.
- D oxygen and water.

32 What are X and Y in the reaction shown?



	X	Y
A	hydrochloric acid	ammonia
B	hydrochloric acid	chlorine
C	sodium hydroxide	ammonia
D	sodium hydroxide	chlorine

33 Which process does **not** produce carbon dioxide?

- A combustion of petrol
- B reaction between an acid and a metal hydroxide
- C reaction between an acid and a carbonate
- D respiration

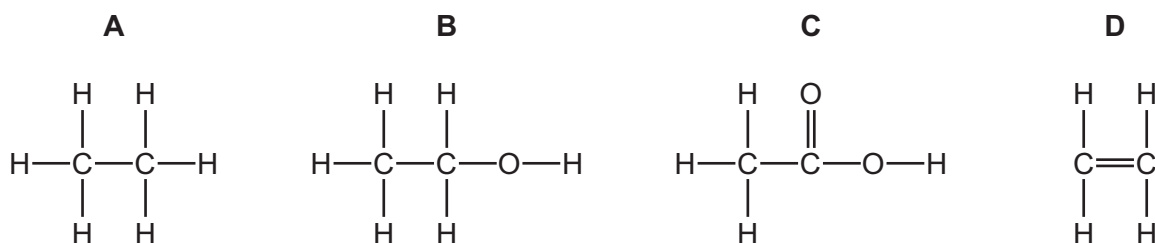
34 Two industrial processes that involve heating are

- extracting iron from its ore using a blast furnace,
- making lime.

In which of these processes is calcium carbonate used?

	extracting iron	making lime
A	✓	✓
B	✓	x
C	x	✓
D	x	x

35 Which structure shows a carboxylic acid?



36 Which method is used to obtain petrol from petroleum?

- A crystallisation
- B diffusion
- C filtration
- D fractional distillation

37 Which group of compounds is part of a homologous series?

- A CH_4 , C_2H_4 , C_3H_8
- B C_3H_6 , C_3H_8 , $\text{C}_3\text{H}_7\text{OH}$
- C CH_3OH , $\text{C}_2\text{H}_5\text{OH}$, $\text{C}_3\text{H}_7\text{OH}$
- D $\text{CH}_3\text{CO}_2\text{H}$, $\text{CH}_3\text{CH}_2\text{OH}$, HCO_2H

38 Bitumen is a substance obtained from the fractional distillation of petroleum.

Which row describes its boiling point and the size of its molecules?

	boiling point	size of molecules
A	high	large
B	high	small
C	low	large
D	low	small

39 A hydrocarbon A is cracked to make B and hydrogen.

Compound C is formed by the addition polymerisation of B.

To which homologous series do A, B and C belong?

	alkene	alkane
A	A	B and C
B	B	A and C
C	C	A and B
D	–	A and C

40 Which statement about ethanol is **not** correct?

- A It can be used as a fuel.
- B It can be used as a solvent.
- C It is formed by reaction between steam and ethane.
- D Its functional group is $-\text{OH}$.

DATA SHEET
The Periodic Table of the Elements

		Group																																										
I	II	III	IV	V	VI	VII	0																																					
		1 H Hydrogen 1											4 He Helium 2																															
7 Li Lithium 3	9 Be Beryllium 4											20 Ne Neon 10																																
23 Na Sodium 11	24 Mg Magnesium 12	27 Al Aluminium 13	28 Si Silicon 14	31 P Phosphorus 15	32 S Sulfur 16	35.5 Cl Chlorine 17	40 Ar Argon 18																																					
39 K Potassium 19	40 Ca Calcium 20	56 Fe Iron 26	55 Mn Manganese 25	59 Co Cobalt 27	59 Ni Nickel 28	64 Cu Copper 29	70 Ga Gallium 31	73 Ge Germanium 32	75 As Arsenic 33	79 Se Selenium 34	80 Br Bromine 35	84 Kr Krypton 36																																
85 Rb Rubidium 37	88 Sr Strontium 38	101 Ru Ruthenium 44	101 Rh Rhodium 45	103 Rh Rhodium 45	106 Pd Palladium 46	108 Ag Silver 47	112 Cd Cadmium 48	115 In Indium 49	119 Sn Tin 50	122 Sb Antimony 51	127 I Iodine 53	131 Xe Xenon 54																																
133 Cs Caesium 55	137 Ba Barium 56	186 Os Osmium 76	186 Re Rhenium 75	184 W Tungsten 74	192 Ir Iridium 77	195 Pt Platinum 78	201 Hg Mercury 80	204 Tl Thallium 81	207 Pb Lead 82	209 Bi Bismuth 83	210 Po Polonium 84	210 Rn Radon 86																																
226 Fr Francium 87	226 Ra Radium 88	227 Ac Actinium 89																																										
*58-71 Lanthanoid series																																												
†90-103 Actinoid series																																												
<table style="width: 100%; border-collapse: collapse;"> <tr> <td style="width: 5%; text-align: center;">a</td> <td style="width: 5%; text-align: center;">X</td> <td style="width: 5%; text-align: center;">b</td> <td style="width: 5%;"></td> <td style="width: 5%; text-align: center;">140 Ce Cerium 58</td> <td style="width: 5%; text-align: center;">141 Pr Praseodymium 59</td> <td style="width: 5%; text-align: center;">144 Nd Neodymium 60</td> <td style="width: 5%; text-align: center;">150 Sm Samarium 62</td> <td style="width: 5%; text-align: center;">152 Eu Europium 63</td> <td style="width: 5%; text-align: center;">157 Gd Gadolinium 64</td> <td style="width: 5%; text-align: center;">162 Dy Dysprosium 66</td> <td style="width: 5%; text-align: center;">165 Ho Holmium 67</td> <td style="width: 5%; text-align: center;">169 Tm Thulium 69</td> <td style="width: 5%; text-align: center;">173 Yb Ytterbium 70</td> <td style="width: 5%; text-align: center;">175 Lu Lutetium 71</td> </tr> <tr> <td style="text-align: center;">Key</td> <td></td> <td></td> <td></td> <td style="text-align: center;">232 Th Thorium 90</td> <td style="text-align: center;">238 Pa Protactinium 91</td> <td style="text-align: center;">238 U Uranium 92</td> <td style="text-align: center;">238 Np Neptunium 93</td> <td style="text-align: center;">238 Am Americium 95</td> <td style="text-align: center;">238 Cm Curium 96</td> <td style="text-align: center;">238 Bk Berkelium 97</td> <td style="text-align: center;">238 Cf Californium 98</td> <td style="text-align: center;">238 Es Einsteinium 99</td> <td style="text-align: center;">238 Fm Fermium 100</td> <td style="text-align: center;">238 Md Mendelevium 101</td> <td style="text-align: center;">238 No Nobelium 102</td> <td style="text-align: center;">238 Lr Lawrencium 103</td> </tr> </table>													a	X	b		140 Ce Cerium 58	141 Pr Praseodymium 59	144 Nd Neodymium 60	150 Sm Samarium 62	152 Eu Europium 63	157 Gd Gadolinium 64	162 Dy Dysprosium 66	165 Ho Holmium 67	169 Tm Thulium 69	173 Yb Ytterbium 70	175 Lu Lutetium 71	Key				232 Th Thorium 90	238 Pa Protactinium 91	238 U Uranium 92	238 Np Neptunium 93	238 Am Americium 95	238 Cm Curium 96	238 Bk Berkelium 97	238 Cf Californium 98	238 Es Einsteinium 99	238 Fm Fermium 100	238 Md Mendelevium 101	238 No Nobelium 102	238 Lr Lawrencium 103
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<p>a = relative atomic mass X = atomic symbol b = proton (atomic) number</p>																																												

The volume of one mole of any gas is 24 dm³ at room temperature and pressure (r.t.p.).

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